Chemistry 212

SPRING 2015

Welcome to Chemistry 212 for the Spring Semester of 2015. This course is the continuation of Chemistry 211 and will cover many basic principles of chemistry. Topics this semester will include: colligative properties, kinetics, equilibrium, acids and bases, solubility, thermodynamics, oxidation and reduction, nuclear chemistry and an introduction to inorganic and bio- chemistry.

Course Title/Number	Principles of Chemistry II - CHM 212
Semester /Year	Spring 2015
Days/Time	MW, 1600p - 1715p
Location	S 473
Instructor	Price, William
Office	S 482
Phone	696-3156
E-Mail	pricew@marshall.edu
Office Hours	MW 15:30 -1600, TR 12:00-1:00
University Policies	By enrolling in this course, you agree to the University Policies listed below. Please read the full text of each policy by going to www.marshall.edu/academic-affairs and clicking on "Marshall University Policies." Or, you can access the policies directly by going to www.marshall.edu/academic-affairs/policies/. Academic Dishonesty/Excused Absence Policy for Undergraduates/Computing Services Acceptable Use/Inclement Weather/Dead Week/Students with Disabilities/Academic Forgiveness/Academic Probation and Suspension/Academic Rights and Responsibilities of Students/Affirmative Action/Sexual Harassment

Course Description: From Catalog

Principles of Chemistry II. 3 hrs. I, II, S. A continuation of Chemistry 211 with emphasis on the inorganic chemistry of the representative elements and transition metals. 3 lec. (PR: *C* or better in CHM 211; CR: CHM 218)

The table below shows the following relationships: How each student learning outcome will be practiced and assessed in the course.

Course student learning outcomes	How students will practice each outcome in this course	How student achievement of each outcome will be assessed in this course
Students will To become familiar with the vocabulary of modern chemistry	Lecture, quizzes, practice assignments	Exams and quizzes
Students will gain insight into the ever-expanding role of chemistry within the context of society, medicine, materials and environment.	Lecture, quizzes, practice assignments	Exams and quizzes

Students will learn and reinforce logical strategies for solving quantitative problems.	Lecture, quizzes, practice assignments	Exams and quizzes

Required Texts, Additional Reading, and Other Materials

Gilbert, Kirss, Foster, Davies *Chemistry* 3rd Edition.
Smartworks KEY : CHEM47772 at http://smartwork.wwnorton.com

Electronic Device Policy

All cell phones and pagers must be either turned off or onto vibrate mode during class. Laptops must be turned off and placed on the floor during the lecture period. During examinations, all electronic devices except calculators must be inaccessible. Students **MUST BRING A CALCULATOR** to class for all lectures and exams. Calculators that are part of a cell phone or smart phones are **not** acceptable for use during an exam or quiz.

Grading Policy

There will be approximately11 quizzes, four midterm exams (one midterm will be dropped) and one cumulative final exam. Quizzes will count for 20 points (one quiz will be dropped), midterm exams will constitute a total of 60 points, while the final exam will be worth 20 points. **BONUS:** completion of *SmartWork* homework can result in an extra 2.5 points. Exam and quiz material will be drawn from the homework, the lecture, and the text. See schedule of tentative exam dates. Missed exams or quizzes may be made up, with a valid University excuse, on **Wednesday May 6th from 12 pm until 2 pm**. If you are planning on making up work on this make-up day you **must** let me know via email by 4 pm Friday, April 24, 2015.
Cutoffs for grades will be no higher than those listed below, but may be lowered if appropriate.

A ≥ 90.00; 90.00 < **B** ≥ 80.00; 80.00 < **C** ≥ 70.00; 70.00 < **D** ≥ 60.00; 60.00 < **F**

Attendance Policy

I strongly encourage you to come to class so that you can more fully understand the material that you will read in the book. If you are absent, obtain the notes from another student or online. In situations where the student is aware of the absence for an exam or quiz in advance, arrangement for accommodations must be made prior to the absence. Otherwise, the designation of an absence as excused and any accommodation for that absence will be decided by the Dean of Students. If a student decides to not complete the course, he or she must visit the registrar and complete the appropriate paperwork to remove the course from his or her schedule. The last day to withdraw from a single class is **Friday, March 27, 2015**.

* Reading assignments and exam dates are approximate and may be subject to change

Tentative Course Schedule*

To make the most of each class period, reading and assignments should be completed before lecture.

Date	Reading	Notes	Week	Reading	Notes
1: 1/12-1/14	*Chapter 10* *Chapter 11*	Review (Forces) Solutions	**9:** 3/11	*Chapters 15 & 16*	Discuss Midterm II
2: 1/21	*Chapter 11*	Solutions	**10:** 3/23-3/25	*Chapter 17*	Equilibrium II Drop Deadline(3/27)
3: 1/26-1/28	*Chapter 14*	Thermodynamics	**11:** 3/30	*Chapter 17*	Equilibrium II
4: 2/2 -2/4	*Chapter 14*	Thermodynamics	**11:** 4/1	*Chapter 18*	Coordination Chemistry
5: 2/9	*Chapters 10, 11 & 14*	Midterm I	**12:** 4/6	*Chapter 18*	Coordination Chemistry
5: 2/11	*Chapters 10, 11 & 14*	Discuss Midterm I	**12:** 4/8	*Chapters 17 & 18*	Midterm 3
6: 2/16-2/18	*Chapter 15*	Chemical Kinetics	**13:** 4/13-4/15	*Chapter 19*	ReDox Rxns
7: 2/23	*Chapter 15*	Chemical Kinetics	**14:** 4/20-4/22	*Chapter19* *Chapter21*	ReDox Rxns Nuclear Chemistry
7: 2/25	*Chapter 16*	Equilibrium	**15:** 4/27	*Chapters 19 & 21*	Midterm 4 (optional)
8: 3/2-3/4	*Chapter 16*	Equilibrium	**15:** 4/29	*Review*	
9: 3/9	*Chapters 15 & 16*	Midterm 2	**15:** 5/2	*Saturday 9:50-11:50*	**Final Exam**

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