

Course Syllabus Spring 2017 Chemistry 205: General Chemistry I

Department of Chemistry, Marshall University

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Office Hours: Study Hall or by appointment

Credit Hours: 3.00 hours

Course Time and Location: 10:00-11:35 M-F; Room 203

Required Materials: Textbook – The Basics of General, Organic, and Biological Chemistry, Spiral Notebook or 3-Ring Binder; Composition Notebook for labs; Calculator; Pen/Pencil

Determination of Course Grade: There will be five exams composed of multiple choice and possibly free response questions. The tests will make up approximately 60% of the final grade. The final exam will count as 20% of the final grade. The remaining 20% will be primarily quizzes. The grading scale for the Marshall University grade will be:

A > 90%, **B** 80 to 89%, **C** 70 to 79%, **D** 60 to 69%, and **F** < 60%. The grading scale for the Chemistry II grade at Poca High School will follow the scale outlined in your handbook for Putnam County Schools.

Attendance: Regular attendance is expected. See me for makeup tests or quizzes.

Homework problems will be assigned for each chapter and will be discussed in class, but will not be collected for a grade. Problems similar to those on the homework will be included on the tests and quizzes. Attendance, reading, and working the homework are essential for successful completion of this course. Plan on reviewing concepts covered in class each day for a few minutes. **Please seek me out if you want or need help.** Please be on time.

Electronic Device Policy: All cell phones and pagers must be turned to vibrate during class. Recording of lectures without the instructor's permission is prohibited. During examinations, all electronic devices except calculators must be inaccessible. Students **MUST BRING A CALCULATOR** to class for all lectures and exams. Calculators that are part of a cell phone or PDA are not acceptable during an exam or quiz.

Chemistry 203 Learning Objectives

Chapter 1

Define chemistry in relation to other sciences.

Identify the general steps in the scientific method.

Use chemical and physical properties, including phase, to describe matter.

Identify a sample of matter as an element, a compound, or a mixture.

Express quantities properly, using a number and a unit.

Express a large number or a small number in scientific notation.
Identify the number of significant figures in a reported value.
Use significant figures correctly in arithmetical operations.
Recognize the SI base units and explain the system of prefixes used with them. Convert a value reported in one unit to a corresponding value a different unit.

Chapter 2

Define a chemical element and give examples of the abundance of different elements.
Represent a chemical element with a chemical symbol.
Explain all matter is composed of atoms.
Describe the modern atomic theory.
Describe the three main subatomic particles.
State how the subatomic particles are arranged in atoms.
Define and differentiate between the atomic number and the mass number of an element.
Explain how isotopes differ from one another.
Define atomic mass and atomic mass unit.
Describe how electrons are grouped within atoms.
Explain how elements are organized into the periodic table.
Describe how some characteristics of elements relate to their positions on the periodic table.

Chapter 3

Define the octet rule.
Describe how ionic bonds are formed.
Define the two types of ions.
Use Lewis diagrams to illustrate ion formation.
Write the chemical formula for a simple ionic compound.
Recognize polyatomic ions in chemical formulas.
Use the rules for naming ionic compounds.
Determine the formula mass of an ionic compound.

Chapter 4

Describe how a covalent bond forms.
Determine the chemical formula of a simple covalent compound from its name.
Determine the name of a simple covalent compound from its chemical formula.
Recognize molecules that are likely to have multiple covalent bonds. Compare covalent bonds in terms of bond length and bond polarity. Determine the molecular mass of a molecule.
Predict the general shape of a simple covalent molecule.
Define *organic chemistry*.
Identify organic molecules as alkanes, alkenes, alkynes, alcohol, or carboxylic acids.

Chapter 5

Correctly define a law as it pertains to science.
State the law of conservation of matter.

Define *chemical reaction*.

Use a balanced chemical equation to represent a chemical reaction.

Calculate the amount of one substance that will react with or be produced from a given amount of another substance.

Classify a given chemical reaction into a variety of types.

Identify a chemical reaction as an oxidation-reduction reaction.

Identify oxidation-reduction reactions with organic compounds.

Chapter 6

Define the mole unit.

Learn how the masses of moles of atoms and molecules are expressed.

Convert quantities between mass units and mole units.

Use a balanced chemical reaction to determine molar relationships between the substances.

Convert from mass or moles of one substance to mass or moles of another substance in a chemical reaction.

Chapter 7

Define *energy* and *heat*.

Relate heat transfer to temperature change.

Determine the heat associated with a phase change.

Define *bond energy*.

Determine if a chemical process is exothermic or endothermic.

Relate the concept of energy change to chemical reactions that occur in the body.

Chapter 8

Define *phase*.

Identify the type of interactions between molecules.

Describe the solid and liquid phases.

Describe the gas phase.

Predict the properties of gases using the gas laws.

Chapter 9

Understand what causes solutions to form.

Express the amount of solute in a solution in various concentration units.

Use molarity to determine quantities in chemical reactions.

Determine the resulting concentration of a diluted solution.

Describe the dissolution process at the molecular level.

Describe how the properties of solutions differ from those of pure solvents.

Chapter 10

Recognize a compound as an Arrhenius acid or an Arrhenius base.

Recognize a compound as a Brønsted-Lowry acid or a Brønsted-Lowry base.

Illustrate the proton transfer process that defines a Brønsted-Lowry acid-base reaction.
Write chemical equations for water acting as an acid and as a base.
Describe the difference between strong and weak acids and bases.
Describe how a chemical reaction reaches chemical equilibrium.
Define the pH scale and use it to describe acids and bases.
Define *buffer* and describe how it reacts with an acid or a base.

Chapter 11

Define and give examples of the major types of radioactivity.
Define *half-life*.
Determine the amount of radioactive substance remaining after a given number of half-lives.
Express amounts of radioactivity in a variety of units.
Learn some applications of radioactivity.
Explain where nuclear energy comes from.
Describe the difference between fission and fusion.

Chapter 12

Recognize the composition and properties typical of organic and inorganic compounds. Identify and name simple (straight-chain) alkanes given formulas and write formulas for straightchain alkanes given their names.
Learn how alkane molecules can have branched chains and recognize compounds that are isomers.
Write condensed structural formulas for alkanes given complete structural formulas.
Draw line-angle formulas given structural formulas.
Name alkanes by the IUPAC system and write formulas for alkanes given IUPAC names.
Identify the physical properties of alkanes and describe trends in these properties.
Identify the main chemical properties of alkanes.
Name halogenated hydrocarbons given formulas and write formulas for these compounds given names.
Name cycloalkanes given their formulas and write formulas for these compounds given their names.

Chapter 13

Name alkenes given formulas and write formulas for alkenes given names.
Recognize that alkenes that can exist as cis-trans isomers.
Classify isomers as cis or trans.
Draw structures for cis-trans isomers given their names.
Identify the physical properties of alkenes and describe trends in these properties.
Write equations for the addition reactions of alkenes with hydrogen, halogens, and water. Draw structures for monomers that can undergo addition polymerization and for four-monomerunit sections of an addition polymer.
Describe the general physical and chemical properties of alkynes.
Name alkynes given formulas and write formulas for alkynes given names.

Describe the bonding in benzene and the way typical reactions of benzene differ from those of the alkenes.

Recognize aromatic compounds from structural formulas.

Name aromatic compounds given formulas.

Write formulas for aromatic compounds given their names.

Chapter 14

Describe functional groups and explain why they are useful in the study of organic chemistry.

Identify the general structure for an alcohol.

Identify the structural feature that classifies alcohols as primary, secondary, or tertiary.

Name alcohols with both common names and IUPAC names.

Explain why the boiling points of alcohols are higher than those of ethers and alkanes of similar molar masses.

Explain why alcohols and ethers of four or fewer carbon atoms are soluble in water while comparable alkanes are not soluble.

Describe how to prepare alcohols from alkenes.

Give two major types of reactions of alcohols.

Describe the result of the oxidation of a primary alcohol.

Describe the result of the oxidation of a secondary alcohol.

Describe the structure and uses of some common polyhydric alcohols.

Describe the structure and uses of some phenols.

Describe the structure difference between an alcohol and an ether that affects physical characteristics and reactivity of each.

Name simple ethers.

Describe the structure and uses of some ethers.

Identify the general structure for an aldehyde and a ketone.

Use common names to name aldehydes and ketones.

Use the IUPAC system to name aldehydes and ketones.

Explain why the boiling points of aldehydes and ketones are higher than those of ethers and alkanes of similar molar masses but lower than those of comparable alcohols.

Compare the solubilities in water of aldehydes and ketones of four or fewer carbon atoms with the solubilities of comparable alkanes and alcohols.

Name the typical reactions that take place with aldehydes and ketones.

Describe some of the uses of common aldehydes and ketones.

Identify thiols (mercaptans) by the presence of an SH group. The mild oxidation of thiols gives disulfides.

Chapter 15

Identify the general structure for a carboxylic acid, an ester, an amine, and an amide.

Identify the functional group for a carboxylic acid, an ester, an amine, and an amide.

Name carboxylic acids with common names.

Name carboxylic acids according to IUPAC nomenclature.

Describe the preparation of carboxylic acids.

Compare the boiling points of carboxylic acids with alcohols of similar molar mass.
Compare the solubilities of carboxylic acids in water with the solubilities of comparable alkanes and alcohols in water.
Name the typical reactions that take place with carboxylic acids.
Describe how carboxylic acids react with basic compounds.
Identify the general structure for an ester.
Use common names to name esters.
Name esters according to the IUPAC system.
Compare the boiling points of esters with alcohols of similar molar mass.
Compare the solubilities of esters in water with the solubilities of comparable alkanes and alcohols in water.
Identify and describe the substances from which most esters are prepared.
Describe the typical reaction that takes place with esters.
Identify the products of an acidic hydrolysis of an ester.
Identify the products of a basic hydrolysis of an ester.
Describe phosphate esters.
Understand why phosphate esters are important in living cells.
Identify the general structure for an amine.
Identify the functional group for amines.
Determine the structural feature that classifies amines as primary, secondary, or tertiary.
Use nomenclature systems to name amines.
Explain why the boiling points of primary and secondary amines are higher than those of alkanes or ethers of similar molar mass but are lower than those of alcohols.
Compare the boiling points of tertiary amines with alcohols, alkanes, and ethers of similar molar mass.
Compare the solubilities in water of amines of five or fewer carbon atoms with the solubilities of comparable alkanes and alcohols in water.
Name the typical reactions that take place with amines.
Describe heterocyclic amines.
Identify the general structure for an amide.
Identify the functional group for an amide.
Name amides with common names.
Name amides according to the IUPAC system.
Compare the boiling points of amides with alcohols of similar molar mass.
Compare the solubilities in water of amides of five or fewer carbon atoms with the solubilities of comparable alkanes and alcohols in water.
Describe the preparation procedure for amides.
Identify the typical reaction that amides undergo.

Chapter 16

Recognize carbohydrates and classify them as mono-, di-, or polysaccharides.
Classify monosaccharides as aldoses or ketoses and as trioses, tetroses, pentoses, or hexoses.
Distinguish between a D sugar and an L sugar.

Identify the structures of D-glucose, D-galactose, and D-fructose and describe how they differ from each other.

Define what is meant by anomers and describe how they are formed.

Explain what is meant by mutarotation.

Identify the physical and chemical properties of monosaccharides.

Identify the structures of sucrose, lactose, and maltose.

Identify the monosaccharides that are needed to form sucrose, lactose, and maltose. Compare and contrast the structures and uses of starch, glycogen, and cellulose.

Chapter 17

Recognize the structures of common fatty acids and classify them as saturated, monounsaturated, or polyunsaturated.

Explain why fats and oils are referred to as triglycerides.

Explain how the fatty acid composition of the triglycerides determines whether a substance is a fat or oil.

Describe the importance of key reactions of triglycerides, such as hydrolysis, hydrogenation, and oxidation.

Identify the distinguishing characteristics of membrane lipids.

Describe membrane components and how they are arranged. Identify the functions of steroids produced in mammals.

Chapter 18

Recognize amino acids and classify them based on the characteristics of their side chains.

Explain how an amino acid can act as both an acid and a base.

Explain how a peptide is formed from individual amino acids.

Explain why the sequence of amino acids in a protein is important.

Describe the four levels of protein structure.

Identify the types of attractive interactions that hold proteins in their most stable threedimensional structure.

Explain what happens when proteins are denatured.

Identify how a protein can be denatured.

Explain the functions of enzymes.

Explain how enzymes are classified and named.

Describe the interaction between an enzyme and its substrate.

Describe how pH, temperature, and the concentration of an enzyme and its substrate influence enzyme activity.

Explain what an enzyme inhibitor is.

Distinguish between reversible and irreversible inhibitors.

Distinguish between competitive and noncompetitive inhibitors. Explain why vitamins are necessary in the diet.

Chapter 19

Identify the different molecules that combine to form nucleotides.
Identify the two types of nucleic acids and the function of each type.
Describe how nucleotides are linked together to form nucleic acids.
Describe the secondary structure of DNA and the importance of complementary base pairing.
Describe how a new copy of DNA is synthesized.
Describe how RNA is synthesized from DNA.
Identify the different types of RNA and the function of each type of RNA.
Describe the characteristics of the genetic code.
Describe how a protein is synthesized from mRNA.
Describe the causes of genetic mutations and how they lead to genetic diseases.
Explain how viruses reproduce in cells.

Topics for CHM 203

Chapter 1

01-Introduction
02-Scientific Notation 03-Significant
Figures
04-Measurements and Units
05-Unit Conversions 06-Density

Chapter 2

07-The Elements
08-Atomic Structure
09-Nuclei of Atoms
10-Arrangement of Electrons
11-Periodic Table

Chapter 3

12-Ionic Bonding
13-Formulas of Ionic Compounds
14-Naming Ions and Ionic Compounds

Chapter 4 (Test on 1-26-17)

15-Covalent Compounds
16-Multiple Covalent Bonds
17-Covalent Bond Characteristics
18-Formula Mass, Molecular Mass
19-Introduction to Organic Chemistry

Chapter 5

20-Chemical Equations

21-Types of Reactions

22-Redox Reactions

Chapter 6 (Test on 2-10-17)

23-The Mole

24-Stoichiometry-Moles 25-Stoichiometry-Mass

Chapter 7

26-Heat and Energy

27-Phase Changes

28-Heat in Chemical Reactions

Chapter 8 (Test on 3-7-17)

29-Intermolecular Forces

30-Gases and Pressure Units

31-PVT Relationships 32-Ideal

Gas Law

Chapter 9 (Test on 3-17-17)

33-Solutions

34-Concentrations

35-Colligative Properties

Chapter 10 (Test on 3-27-17)

36-Arrhenius Acids and Bases

37-Bronsted-Lowery Acids and Bases

38-Acid and Base Strengths and Buffers

Chapter 11 (Test on 4-3-17)

39-Radioactivity

40-Radiation Units and Half Life

41-Uses of Radiation and Nuclear Chemistry

Chapter 12

42-Alkanes

44-IUPAC Nomenclature

45-Cycloalkanes
46-Chemical and Physical Properties of Alkanes

Chapter 13

46-Alkenes and Alkynes
47-Geometric Isomers
48-Properties of Alkenes and Alkynes
49-Polymers 50-Aromatic
Compounds

Chapter 14 (Test on 4-26-17)

51-Alcohols
52-Formation and Properties of Alcohols
53-Phenols, Ethers, and Organosulfur Compounds
54-Aldehydes and Ketones

Chapter 15

55-Carboxylic Acids
56-Formation and Properties of Carboxylic Acids
57-Esters
58-Formation, Properties, and Reactions of Esters
59-Amines
60-Properties of Amines 61-Amides

Chapter 16

62-Carbohydrates and Stereoisomers
63-Monosaccharides
64-Di- and Polysaccharides

Chapter 17 (Test on 5-1-17)

65-Fatty Acids
66-Fats and Oils
67-Membranes and Membrane Lipids
68-Steroids

Chapter 18

69-Amino Acids
70-The Isoelectric Point
71-Peptides
72-Proteins
73-Properties of Proteins
74-Enzymes

75-Enzyme Activity
76-Enzyme Inhibition

Chapter 19 (Final Exam 5-8-17)

77-Nucleotides
78-Nucleic Acids
79-Replication and Transcription
80-Mutations and Viruses